**Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date: \_\_\_\_\_\_\_\_\_\_\_\_\_\_ Period: \_\_\_\_\_\_\_**

**Titrations Practice Worksheet**

*Find the requested quantities in the following problems:*

1. Write a balanced equation for the reaction between NaOH and HCl. Use this equation to answer problems 2-4
2. If it takes 54 mL of 0.1 M NaOH to neutralize 125 mL of an HCl solution, what is the concentration of the HCl?
3. If it takes 25 mL of 0.05 M HCl to neutralize 345 mL of NaOH solution, what is the concentration of the NaOH solution?
4. A 25.0 mL sample of HCl was titrated to the endpoint with 15.0 mL of 2.0 M NaOH. What is the molarity of HCl?
5. Write a balanced equation for the reaction between Ca(OH)2 and HNO3. Use this equation to answer problems 6 -8.
6. If it takes 50.0 mL of 0.5 M Ca(OH)2 solution to completely neutralize 125 mL of HNO3solution, what is the concentration of the HNO3 solution?
7. How many mL of 0.50 M HNO3 is necessary to titrate 25.0 mL of 0.05 M Ca(OH)2 solution to the endpoint?
8. If it takes 75.0 mL of 1.5 M HNO3 solution to completely neutralize 125 mL of Ca(OH)2what is the concentration of the Ca(OH)2 solution?